

## 2,2,2-Tribromo-*N*-(4-methylphenyl)-acetamide

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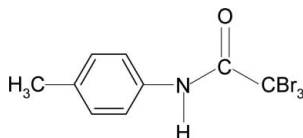
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Key indicators: single-crystal X-ray study;  $T = 299$  K; mean  $\sigma(\text{C}-\text{C}) = 0.015$  Å;  $R$  factor = 0.082;  $wR$  factor = 0.151; data-to-parameter ratio = 16.2.

The asymmetric unit of the title compound,  $\text{C}_9\text{H}_8\text{Br}_3\text{NO}$ , contains two independent molecules which differ in the orientation of the tribromo group. A weak intramolecular  $\text{N}-\text{H}\cdots\text{Br}$  hydrogen bond is observed in each molecule. In the crystal, the independent molecules are linked into chains along the  $b$  axis by intermolecular  $\text{N}-\text{H}\cdots\text{O}$  hydrogen bonds.

### Related literature

For the preparation of the title compound, see: Gowda *et al.* (2003). For our study of the effect of ring and the side-chain substituents on the solid-state structures of *N*-aromatic amides and for similar structures, see: Brown (1966); Gowda *et al.* (2009*a,b,c*).



### Experimental

#### Crystal data

$\text{C}_9\text{H}_8\text{Br}_3\text{NO}$   
 $M_r = 385.89$   
Monoclinic,  $P2_1/c$   
 $a = 9.6926$  (6) Å  
 $b = 20.531$  (1) Å  
 $c = 11.8139$  (8) Å  
 $\beta = 102.664$  (7)°

$V = 2293.8$  (2) Å<sup>3</sup>  
 $Z = 8$   
Mo  $K\alpha$  radiation  
 $\mu = 10.52$  mm<sup>-1</sup>  
 $T = 299$  K  
 $0.48 \times 0.40 \times 0.30$  mm

#### Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector  
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2009)

$T_{\min} = 0.081$ ,  $T_{\max} = 0.145$   
14714 measured reflections  
4121 independent reflections  
3375 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.089$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.082$   
 $wR(F^2) = 0.151$   
 $S = 1.28$   
4121 reflections

255 parameters  
H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 1.34$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.88$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1N}\cdots\text{O2}$	0.86	2.27	3.078 (10)	156
$\text{N1}-\text{H1N}\cdots\text{Br2}$	0.86	2.61	3.111 (8)	118
$\text{N2}-\text{H2N}\cdots\text{O1}^1$	0.86	2.27	3.032 (10)	148
$\text{N2}-\text{H2N}\cdots\text{Br4}$	0.86	2.56	3.051 (9)	118

Symmetry code: (i)  $-x + 1, y - \frac{1}{2}, -z + \frac{3}{2}$ .

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2009); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI5053).

### References

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**supplementary materials**

*Acta Cryst.* (2010). E66, o849 [ doi:10.1107/S1600536810009177 ]

## 2,2,2-Tribromo-*N*-(4-methylphenyl)acetamide

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### Comment

As part of a study of the effect of the ring and the side chain substituents on solid state structures of *N*-aromatic amides (Gowda *et al.*, 2009*a,b,c*), in the present work, the crystal structure of 2,2,2-tribromo-*N*-(4-methylphenyl)acetamide has been determined (Fig. 1). The asymmetric unit of the structure contains two independent molecules, which differ in the orientation of the tribromo group as is evident from either the C-N-CO-CBr<sub>3</sub> or N-CO-C-Br torsional angles. The conformations of the N—H bonds in both molecules are *anti* to the C=O bonds in the side chains, similar to those observed in 2,2,2-tribromo-*N*-(3-methylphenyl)acetamide (Gowda *et al.*, 2009*a*), 2,2,2-tribromo-*N*-(phenyl)acetamide (Gowda *et al.*, 2009*b*), 2,2,2-tribromo-*N*-(4-chlorophenyl)acetamide (Gowda *et al.*, 2009*c*) and other amides (Brown, 1966). The structure of the title compound shows both the intramolecular N—H $\cdots$ Br and intermolecular N—H $\cdots$ O hydrogen bonding.

The packing diagram of molecules showing the hydrogen bonds N1—H1N $\cdots$ O2 and N2—H2N $\cdots$ O1 (Table 1) involved in the formation of molecular chains is shown in Fig. 2.

### Experimental

The title compound was prepared from *p*-toluidine, tribromoacetic acid and phosphorylchloride according to the literature method (Gowda *et al.*, 2003). The purity of the compound was checked by determining its melting point. It was further characterized by recording its infrared spectra. Single crystals of the title compound used for X-ray diffraction studies were obtained by a slow evaporation of its solution in petroleum ether at room temperature.

### Refinement

H atoms were positioned with idealized geometry using a riding model with C—H = 0.93–0.96 Å, N—H = 0.86 Å and  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{parent atom})$ . The residual electron-density features are located in the region of Br4 and Br3. The highest peak is 0.99 Å from Br4 and the deepest hole is 1.28 Å from Br3. Owing to the poor diffraction quality of the crystal, the  $R_{\text{int}}$  value is high (0.089) and this is a structure of relatively low precision.

### Figures

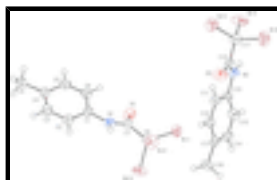


Fig. 1. The asymmetric unit of the title compound, showing the atom labelling scheme. Displacement ellipsoids are drawn at the 50% probability level and H atoms are represented as small spheres of arbitrary radii.

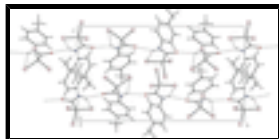


Fig. 2. Molecular packing in the crystal structure of the title compound, with hydrogen bonds shown as dashed lines.

## 2,2,2-Tribromo-*N*-(4-methylphenyl)acetamide

### Crystal data

$C_9H_8Br_3NO$	$F(000) = 1456$
$M_r = 385.89$	$D_x = 2.235 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/c$	Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
Hall symbol: $-P 2_1/c$	Cell parameters from 5162 reflections
$a = 9.6926 (6) \text{ \AA}$	$\theta = 2.7\text{--}27.8^\circ$
$b = 20.531 (1) \text{ \AA}$	$\mu = 10.52 \text{ mm}^{-1}$
$c = 11.8139 (8) \text{ \AA}$	$T = 299 \text{ K}$
$\beta = 102.664 (7)^\circ$	Prism, colourless
$V = 2293.8 (2) \text{ \AA}^3$	$0.48 \times 0.40 \times 0.30 \text{ mm}$
$Z = 8$	

### Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector	4121 independent reflections
Radiation source: fine-focus sealed tube graphite	3375 reflections with $I > 2\sigma(I)$
Rotation method data acquisition using $\omega$ and $\varphi$ scans.	$R_{\text{int}} = 0.089$
Absorption correction: multi-scan (CrysAlis RED; Oxford Diffraction, 2009)	$\theta_{\text{max}} = 25.4^\circ$ , $\theta_{\text{min}} = 4.2^\circ$
$T_{\text{min}} = 0.081$ , $T_{\text{max}} = 0.145$	$h = -11 \rightarrow 11$
14714 measured reflections	$k = -21 \rightarrow 24$
	$l = -13 \rightarrow 14$

### Refinement

Refinement on $F^2$	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.082$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.151$	H-atom parameters constrained
$S = 1.28$	$w = 1/[\sigma^2(F_o^2) + (0.P)^2 + 26.8682P]$
4121 reflections	where $P = (F_o^2 + 2F_c^2)/3$
255 parameters	$(\Delta/\sigma)_{\text{max}} = 0.001$
0 restraints	$\Delta\rho_{\text{max}} = 1.34 \text{ e \AA}^{-3}$
	$\Delta\rho_{\text{min}} = -0.88 \text{ e \AA}^{-3}$

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.7165 (11)	0.2486 (4)	0.6954 (9)	0.029 (2)
C2	0.7414 (11)	0.1962 (5)	0.6292 (10)	0.035 (3)
H2	0.6948	0.1569	0.6335	0.043*
C3	0.8348 (12)	0.2017 (5)	0.5568 (11)	0.043 (3)
H3	0.8516	0.1658	0.5136	0.051*
C4	0.9039 (11)	0.2598 (5)	0.5475 (9)	0.034 (2)
C5	0.8797 (12)	0.3110 (5)	0.6175 (10)	0.039 (3)
H5	0.9279	0.3500	0.6149	0.047*
C6	0.7875 (12)	0.3065 (5)	0.6901 (10)	0.037 (3)
H6	0.7730	0.3419	0.7353	0.044*
C7	0.5389 (11)	0.2876 (4)	0.7992 (9)	0.028 (2)
C8	0.4213 (11)	0.2643 (4)	0.8592 (9)	0.029 (2)
C9	1.0009 (12)	0.2673 (6)	0.4652 (10)	0.043 (3)
H9A	0.9464	0.2667	0.3869	0.052*
H9B	1.0675	0.2320	0.4762	0.052*
H9C	1.0507	0.3079	0.4800	0.052*
Br1	0.37301 (14)	0.33225 (5)	0.95617 (12)	0.0480 (4)
Br2	0.46539 (15)	0.18686 (5)	0.95318 (12)	0.0510 (4)
Br3	0.25450 (14)	0.24690 (6)	0.73598 (12)	0.0531 (4)
N1	0.6157 (9)	0.2404 (4)	0.7652 (8)	0.033 (2)
H1N	0.6029	0.2015	0.7877	0.039*
O1	0.5509 (8)	0.3449 (3)	0.7822 (7)	0.042 (2)
C10	0.3798 (10)	0.0087 (4)	0.6530 (9)	0.025 (2)
C11	0.3488 (11)	-0.0253 (5)	0.5503 (10)	0.036 (3)
H11	0.4045	-0.0606	0.5390	0.044*
C12	0.2340 (11)	-0.0068 (5)	0.4633 (10)	0.034 (3)
H12	0.2144	-0.0297	0.3937	0.041*
C13	0.1478 (10)	0.0452 (5)	0.4785 (9)	0.029 (2)
C14	0.1792 (12)	0.0765 (5)	0.5829 (11)	0.038 (3)
H14	0.1220	0.1108	0.5960	0.046*
C15	0.2935 (12)	0.0590 (5)	0.6701 (10)	0.035 (3)
H15	0.3117	0.0813	0.7403	0.042*
C16	0.5967 (10)	0.0331 (5)	0.7916 (10)	0.030 (2)

## supplementary materials

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C17	0.7240 (12)	0.0074 (5)	0.8818 (10)	0.036 (3)
C18	0.0294 (13)	0.0685 (5)	0.3837 (10)	0.042 (3)
H18A	0.0032	0.0346	0.3270	0.050*
H18B	0.0595	0.1062	0.3475	0.050*
H18C	-0.0504	0.0795	0.4157	0.050*
Br4	0.75931 (14)	-0.08465 (6)	0.87190 (14)	0.0582 (4)
Br5	0.89309 (14)	0.05436 (7)	0.86845 (12)	0.0540 (4)
Br6	0.69036 (17)	0.02579 (8)	1.03563 (12)	0.0661 (4)
N2	0.5023 (9)	-0.0097 (4)	0.7398 (8)	0.032 (2)
H2N	0.5144	-0.0501	0.7586	0.038*
O2	0.5901 (9)	0.0912 (3)	0.7746 (8)	0.049 (2)

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.024 (5)	0.027 (5)	0.037 (7)	0.003 (4)	0.006 (5)	0.004 (4)
C2	0.032 (6)	0.025 (5)	0.049 (8)	0.002 (4)	0.007 (6)	-0.001 (5)
C3	0.039 (7)	0.034 (6)	0.057 (9)	0.005 (5)	0.015 (6)	-0.011 (5)
C4	0.031 (6)	0.045 (6)	0.020 (6)	0.008 (5)	-0.004 (5)	0.007 (5)
C5	0.032 (6)	0.036 (6)	0.049 (8)	-0.008 (5)	0.006 (6)	-0.003 (5)
C6	0.035 (6)	0.033 (5)	0.043 (7)	-0.006 (5)	0.008 (6)	-0.009 (5)
C7	0.030 (6)	0.027 (5)	0.028 (6)	0.001 (4)	0.006 (5)	0.001 (4)
C8	0.030 (6)	0.022 (5)	0.029 (6)	0.005 (4)	-0.008 (5)	-0.002 (4)
C9	0.035 (7)	0.055 (7)	0.035 (7)	-0.003 (6)	-0.004 (6)	-0.009 (6)
Br1	0.0487 (7)	0.0402 (6)	0.0610 (9)	-0.0018 (5)	0.0251 (7)	-0.0155 (6)
Br2	0.0660 (9)	0.0386 (6)	0.0484 (8)	0.0034 (6)	0.0128 (7)	0.0101 (6)
Br3	0.0423 (7)	0.0586 (8)	0.0496 (9)	-0.0124 (6)	-0.0086 (6)	-0.0035 (6)
N1	0.035 (5)	0.015 (4)	0.049 (6)	0.000 (4)	0.009 (5)	-0.002 (4)
O1	0.042 (5)	0.021 (4)	0.064 (6)	0.004 (3)	0.017 (4)	0.001 (3)
C10	0.024 (5)	0.022 (5)	0.027 (6)	-0.003 (4)	0.001 (5)	0.004 (4)
C11	0.027 (6)	0.030 (5)	0.051 (8)	0.001 (5)	0.007 (6)	-0.002 (5)
C12	0.032 (6)	0.037 (6)	0.032 (7)	0.000 (5)	0.004 (5)	-0.003 (5)
C13	0.024 (5)	0.028 (5)	0.034 (7)	0.000 (4)	0.007 (5)	0.008 (5)
C14	0.035 (6)	0.028 (5)	0.053 (8)	0.007 (5)	0.014 (6)	0.003 (5)
C15	0.048 (7)	0.026 (5)	0.030 (7)	-0.001 (5)	0.002 (6)	-0.010 (5)
C16	0.023 (5)	0.026 (5)	0.044 (7)	-0.001 (4)	0.012 (5)	-0.003 (5)
C17	0.034 (6)	0.026 (5)	0.044 (7)	-0.009 (5)	0.001 (6)	-0.004 (5)
C18	0.045 (7)	0.046 (6)	0.031 (7)	0.006 (6)	0.002 (6)	0.009 (5)
Br4	0.0472 (8)	0.0352 (6)	0.0788 (11)	0.0055 (5)	-0.0149 (7)	0.0032 (6)
Br5	0.0393 (7)	0.0677 (8)	0.0523 (9)	-0.0199 (6)	0.0043 (6)	-0.0008 (7)
Br6	0.0644 (10)	0.0966 (11)	0.0391 (8)	-0.0102 (8)	0.0150 (7)	-0.0027 (7)
N2	0.033 (5)	0.018 (4)	0.040 (6)	-0.002 (4)	-0.001 (4)	-0.001 (4)
O2	0.044 (5)	0.022 (4)	0.076 (7)	-0.008 (3)	-0.001 (5)	0.005 (4)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

C1—C2	1.383 (14)	C10—C15	1.371 (13)
C1—C6	1.382 (13)	C10—C11	1.374 (14)
C1—N1	1.420 (12)	C10—N2	1.439 (13)

C2—C3	1.380 (15)	C11—C12	1.392 (15)
C2—H2	0.93	C11—H11	0.93
C3—C4	1.384 (15)	C12—C13	1.392 (13)
C3—H3	0.93	C12—H12	0.93
C4—C5	1.389 (14)	C13—C14	1.366 (15)
C4—C9	1.501 (15)	C13—C18	1.495 (15)
C5—C6	1.371 (15)	C14—C15	1.384 (16)
C5—H5	0.93	C14—H14	0.93
C6—H6	0.93	C15—H15	0.93
C7—O1	1.201 (11)	C16—O2	1.209 (11)
C7—N1	1.337 (12)	C16—N2	1.318 (13)
C7—C8	1.545 (14)	C16—C17	1.536 (15)
C8—Br1	1.927 (9)	C17—Br4	1.928 (10)
C8—Br2	1.932 (9)	C17—Br5	1.938 (10)
C8—Br3	1.957 (10)	C17—Br6	1.953 (11)
C9—H9A	0.96	C18—H18A	0.96
C9—H9B	0.96	C18—H18B	0.96
C9—H9C	0.96	C18—H18C	0.96
N1—H1N	0.86	N2—H2N	0.86
C2—C1—C6	119.6 (9)	C15—C10—C11	119.3 (10)
C2—C1—N1	117.7 (8)	C15—C10—N2	121.8 (9)
C6—C1—N1	122.8 (9)	C11—C10—N2	118.8 (9)
C3—C2—C1	120.5 (9)	C10—C11—C12	119.9 (10)
C3—C2—H2	119.8	C10—C11—H11	120.1
C1—C2—H2	119.8	C12—C11—H11	120.1
C2—C3—C4	120.9 (10)	C13—C12—C11	121.3 (10)
C2—C3—H3	119.5	C13—C12—H12	119.4
C4—C3—H3	119.5	C11—C12—H12	119.4
C3—C4—C5	117.2 (10)	C14—C13—C12	117.2 (10)
C3—C4—C9	121.5 (10)	C14—C13—C18	120.6 (9)
C5—C4—C9	121.3 (10)	C12—C13—C18	122.2 (10)
C6—C5—C4	122.7 (10)	C13—C14—C15	122.2 (10)
C6—C5—H5	118.7	C13—C14—H14	118.9
C4—C5—H5	118.7	C15—C14—H14	118.9
C5—C6—C1	119.1 (10)	C10—C15—C14	120.1 (10)
C5—C6—H6	120.5	C10—C15—H15	120.0
C1—C6—H6	120.5	C14—C15—H15	120.0
O1—C7—N1	125.4 (9)	O2—C16—N2	125.0 (11)
O1—C7—C8	119.2 (8)	O2—C16—C17	117.4 (9)
N1—C7—C8	115.3 (8)	N2—C16—C17	117.6 (8)
C7—C8—Br1	110.4 (6)	C16—C17—Br4	114.9 (7)
C7—C8—Br2	115.2 (6)	C16—C17—Br5	109.7 (7)
Br1—C8—Br2	107.8 (5)	Br4—C17—Br5	108.5 (5)
C7—C8—Br3	106.7 (7)	C16—C17—Br6	107.9 (7)
Br1—C8—Br3	107.8 (5)	Br4—C17—Br6	108.3 (5)
Br2—C8—Br3	108.7 (5)	Br5—C17—Br6	107.2 (5)
C4—C9—H9A	109.5	C13—C18—H18A	109.5
C4—C9—H9B	109.5	C13—C18—H18B	109.5
H9A—C9—H9B	109.5	H18A—C18—H18B	109.5

## supplementary materials

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C4—C9—H9C	109.5	C13—C18—H18C	109.5
H9A—C9—H9C	109.5	H18A—C18—H18C	109.5
H9B—C9—H9C	109.5	H18B—C18—H18C	109.5
C7—N1—C1	126.0 (8)	C16—N2—C10	122.4 (8)
C7—N1—H1N	117.0	C16—N2—H2N	118.8
C1—N1—H1N	117.0	C10—N2—H2N	118.8
C6—C1—C2—C3	-1.0 (17)	C15—C10—C11—C12	2.5 (14)
N1—C1—C2—C3	178.0 (10)	N2—C10—C11—C12	-177.6 (9)
C1—C2—C3—C4	-1.0 (18)	C10—C11—C12—C13	-0.8 (15)
C2—C3—C4—C5	2.7 (17)	C11—C12—C13—C14	-1.2 (14)
C2—C3—C4—C9	-177.3 (11)	C11—C12—C13—C18	176.1 (10)
C3—C4—C5—C6	-2.6 (17)	C12—C13—C14—C15	1.5 (15)
C9—C4—C5—C6	177.4 (11)	C18—C13—C14—C15	-175.8 (10)
C4—C5—C6—C1	0.7 (18)	C11—C10—C15—C14	-2.2 (15)
C2—C1—C6—C5	1.2 (17)	N2—C10—C15—C14	177.9 (9)
N1—C1—C6—C5	-177.8 (11)	C13—C14—C15—C10	0.1 (16)
O1—C7—C8—Br1	-26.3 (12)	O2—C16—C17—Br4	-162.1 (8)
N1—C7—C8—Br1	156.8 (8)	N2—C16—C17—Br4	18.7 (12)
O1—C7—C8—Br2	-148.7 (9)	O2—C16—C17—Br5	-39.5 (12)
N1—C7—C8—Br2	34.4 (12)	N2—C16—C17—Br5	141.3 (8)
O1—C7—C8—Br3	90.6 (10)	O2—C16—C17—Br6	77.0 (10)
N1—C7—C8—Br3	-86.3 (9)	N2—C16—C17—Br6	-102.2 (9)
O1—C7—N1—C1	-5.5 (18)	O2—C16—N2—C10	1.6 (16)
C8—C7—N1—C1	171.2 (10)	C17—C16—N2—C10	-179.3 (9)
C2—C1—N1—C7	-153.2 (11)	C15—C10—N2—C16	-49.1 (14)
C6—C1—N1—C7	25.8 (17)	C11—C10—N2—C16	130.9 (10)

### Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1N...O2	0.86	2.27	3.078 (10)	156
N1—H1N...Br2	0.86	2.61	3.111 (8)	118
N2—H2N...O1 <sup>i</sup>	0.86	2.27	3.032 (10)	148
N2—H2N...Br4	0.86	2.56	3.051 (9)	118

Symmetry codes: (i)  $-x+1, y-1/2, -z+3/2$ .



Fig. 1

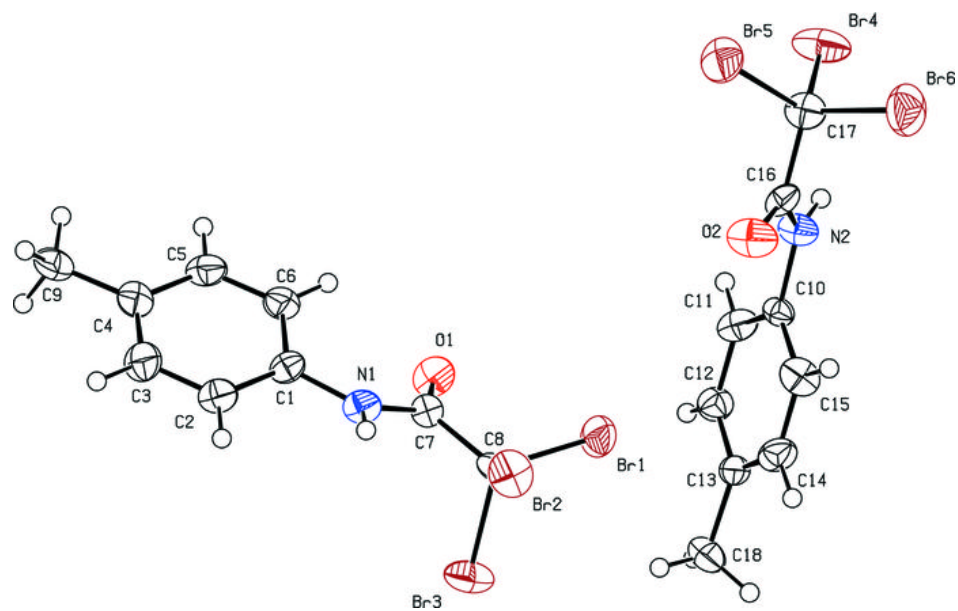


Fig. 2

